• **Nuclide** = a particular type of nucleus, characterized by a specific number of protons and neutrons and therefore a specific atomic number and nucleon number.

• **Nucleon number** or **mass number** = the number of **nucleons** (protons and neutrons) in the nucleus of a nuclide.
Mass number (nucleon number)

Atomic number

Element symbol
Radioactive Iodine

• One of the products of the fission reaction of uranium atoms with 92 protons and 143 neutrons is iodine atoms with 53 protons and 78 neutrons.

\[
\frac{235}{92}U \quad 2^{35}U \quad U-235 \quad \text{uranium-235}
\]

\[
\frac{131}{53}I \quad 1^{31}I \quad I-131 \quad \text{iodine-131}
\]
Two Forces in Nucleus

- **Electromagnetic force** = the force that causes opposite electrical charges to attract each other and like charges to repel each other.

- **Strong force** = the attractive force between nucleons (protons and neutrons).
Formation of a Helium Nucleus

- Helium-2 with just two protons nucleus is unstable.
  \[ p + p \rightarrow _2^2\text{He}^{2+} \]

- The shorter the distance between the protons is, the stronger the electromagnetic repulsion between them.

- When they are close enough to form a helium nucleus, the strong force is not strong enough to overcome the electromagnetic repulsion, so the protons are pushed apart.
Nuclear Stability

- Neutrons increase the attraction from the strong force without increasing electromagnetic repulsion between nucleons.
- Combining two neutrons with two protons increases the strong force enough to overcome the electromagnetic repulsion, making a stable helium nucleus.

\[
p + p + n + n \rightarrow ^4_2\text{He}^{2+} + \text{Energy}
\]
Band of Stability

Island of stability

One to one neutron to proton ratio
Radioactivity

- Too many neutrons (neutron rich)
- Beta emission (beta decay)
- Too few neutrons
  - Positron emission (positron decay)
  - or electron capture
- Too many protons
  - Alpha emission (alpha decay)
Alpha Emission

$\frac{238}{92}\text{U} \rightarrow \frac{234}{90}\text{Th} + \frac{4}{2}\text{He}$

Two protons and two neutrons lost

The protons and neutrons leave as an alpha particle.

Energy
Beta Emission

\[ ^{131}_{53}I \rightarrow ^{131}_{54}Xe + ^0_{-1}e \]

A neutron becomes a proton (which stays in the nucleus) and an electron (which is ejected from the atom).
Positron Emission

$$^{40}_{19}\text{K} \rightarrow ^{40}_{18}\text{Ar} + ^{0}_{+1}\text{e}$$

A proton becomes a neutron (which stays in the nucleus) and a positron (which is ejected from the atom).
Electron Capture

\[ ^{0}_{-1}\text{e} + ^{125}_{53}\text{I} \rightarrow ^{125}_{52}\text{Te} \]

An electron combines with a proton to form a neutron.

\[ e^- \rightarrow \text{proton} + \text{neutron} \]

Energy
$^{137}_{55} \text{Cs} \rightarrow ^{137}_{56} \text{Ba}^* + ^0_{-1} \text{e}$

$\rightarrow ^{137}_{56} \text{Ba} + \gamma$-photon

Gamma Emission

Excited state
Nuclear reactions involve changes in the nucleus, whereas chemical reactions involve the loss, gain, and sharing of electrons.

Different isotopes of the same element may undergo very different nuclear reactions, even though an element’s isotopes all share the same chemical characteristics.
Nuclear Reactions (2)

• Unlike chemical reactions, the rates of nuclear reactions are unaffected by temperature, pressure, and the presence of other atoms to which the radioactive atom may be bonded.

• Nuclear reactions, in general, give off much more energy than chemical reactions.
Nuclear Equations

**Alpha emission**

mass number \( 238 \)  \( \rightarrow \)  \( 234 \)  +  \( 4 \)  =  \( 238 \)

atomic number \( 92 \)  \( \rightarrow \)  \( 90 \)  +  \( 2 \)  =  \( 92 \)

\[ \frac{238}{92} \text{U} \rightarrow \frac{234}{90} \text{Ih} + \frac{4}{2} \text{He} \]

**Beta emission**

mass number \( 131 \)  \( \rightarrow \)  \( 131 \)  +  \( 0 \)  =  \( 131 \)

atomic number \( 53 \)  \( \rightarrow \)  \( 54 \)  +  \( (-1) \)  =  \( 53 \)

\[ \frac{131}{53} \text{I} \rightarrow \frac{131}{54} \text{Xe} + \frac{0}{-1} \text{e} \]

**Positron emission**

mass number \( 40 \)  \( \rightarrow \)  \( 40 \)  +  \( 0 \)  =  \( 40 \)

atomic number \( 19 \)  \( \rightarrow \)  \( 18 \)  +  \( 1 \)  =  \( 19 \)

\[ \frac{40}{19} \text{K} \rightarrow \frac{40}{18} \text{Ar} + \frac{0}{+1} \text{e} \]

**Electron capture**

mass number \( 0 \)  +  \( 125 \)  =  \( 125 \)

atomic number \( -1 \)  +  \( 53 \)  =  \( 52 \)

\[ \frac{0}{-1} \text{e} + \frac{125}{53} \text{I} \rightarrow \frac{125}{52} \text{Te} \]
General Nuclear Equations

**Alpha emission**

\[
\frac{A}{Z} X \longrightarrow \frac{A-4}{Z-2} Y + \frac{4}{2} \text{He}
\]

**Beta emission**

\[
\frac{A}{Z} X \longrightarrow \frac{A}{Z+1} Y + \frac{0}{-1} e
\]

**Positron emission**

\[
\frac{A}{Z} X \longrightarrow \frac{A}{Z-1} Y + \frac{0}{+1} e
\]

**Electron capture**

\[
\frac{0}{-1} e + \frac{A}{Z} X \longrightarrow \frac{A}{Z-1} Y
\]
**Half-life** = the time it takes for one-half of a sample to disappear.
Ionization by Alpha Particles

1. A positively charged alpha particle attracts electrons enough to drag an electron off of an uncharged atom or molecule to form a cation.

2. The electron can combine with another uncharged atom or molecule to form an anion.

3. The high-velocity alpha particle continues on and can create many ions.
Ionization by Beta Particles

1. A negatively charged beta particle repels electrons enough to push an electron off of an uncharged atom or molecule to form a cation.

2. The electron can combine with another uncharged atom or molecule to form an anion.

3. The high-velocity beta particle continues on and can create many ions.
Ionization by Gamma Rays

1. When a gamma ray collides with an uncharged atom or molecule, it excites an electron to such a high energy level that it is removed completely to form a cation.

2. The electron released might be moving fast enough to push electrons off other atoms and molecules to form many ions.

3. The electron released can combine with another uncharged atom or molecule to form an anion.
Radiation Effect on Body

- As the radioactive emissions ionize atoms and molecules, such as water molecules, they also form highly reactive free radicals, which are particles with unpaired electrons.

\[
\begin{align*}
H_2O & \rightarrow H_2O^{+} + e^- \\
H_2O^{+} + H_2O & \rightarrow H_3O^+ + \cdot OH \\
H_2O + e^- & \rightarrow H\cdot + OH^-
\end{align*}
\]

- These reactive particles react with important substances in the body, leading to immediate damage and delayed problems, such as cancer.
Penetration by Radioactive Emissions

- There is an animation that will provide a review of radioactivity at the following web address.
- A portion of this animation describes the relative penetrating ability of alpha particles, beta particles, and gamma photons.

https://preparatorychemistry.com/radioactivity_Canvas.html
Effects on Body from Inside and Outside

- **Internal emitters** = emitters of alpha and weak beta particles
  - Cannot penetrate dead outer layer of skin,
  - Dangerous if inside the body (e.g. Pu-239)
- **External emitters** = emitters of gamma rays, strong beta particles, or neutrons
  - Dangerous from outside the body

http://preparatorychemistry.com/radioactivity.html
Relative Biological Effectiveness

- The ratio of biological effectiveness of one type of ionizing radiation relative to another, given the same amount of absorbed energy.

<table>
<thead>
<tr>
<th>Type</th>
<th>RBE (QF = quality factor)</th>
</tr>
</thead>
<tbody>
<tr>
<td>x-rays</td>
<td>1</td>
</tr>
<tr>
<td>γ-rays</td>
<td>1</td>
</tr>
<tr>
<td>β-particles</td>
<td>1-3</td>
</tr>
<tr>
<td>α-particles</td>
<td>5-20</td>
</tr>
<tr>
<td>neutrons</td>
<td>5-20</td>
</tr>
</tbody>
</table>

http://en.wikipedia.org/wiki/Relative_Biological_Effectiveness
• **Physical dose** = energy absorbed per kg of tissue:
  - 1 rad = 100 erg/g = 0.01 Gy
  - 1 Gray (Gy) = 1 J/kg = 100 rad

• **Biological dose** = physical dose × biological effectiveness:
  - QF = Quality Factor
    = RBE = relative biological effectiveness
  - 1 rem = 1 rad × (QF = 1)
  - ? rem = # rad × QF
  - 1 Sievert (Sv) = 1 Gy × (QF = 1) = 100 rem
  - ? Sv = # Gy × QF
<table>
<thead>
<tr>
<th>$S_v$</th>
<th>Effect</th>
</tr>
</thead>
<tbody>
<tr>
<td>$&gt; 0.25$</td>
<td>change in blood counts</td>
</tr>
<tr>
<td>$&gt; 0.50$</td>
<td>temporary sterility</td>
</tr>
<tr>
<td>1-2</td>
<td>vomiting, hair loss, etc.</td>
</tr>
<tr>
<td>2-3</td>
<td>5-35% fatalities in 30 d</td>
</tr>
<tr>
<td>$&gt; 3$</td>
<td>permanent sterility</td>
</tr>
<tr>
<td>4-5</td>
<td>50% fatalities in 30 d ($LD_{50/30}$)</td>
</tr>
<tr>
<td>6-8</td>
<td>95% fatalities in 30 d</td>
</tr>
<tr>
<td>$&gt; 10$</td>
<td>death in 10 d</td>
</tr>
<tr>
<td>$&gt; 50$</td>
<td>death in 2 d</td>
</tr>
<tr>
<td>$&gt; 100$</td>
<td>immediate death</td>
</tr>
</tbody>
</table>
# Acute Dose Examples

<table>
<thead>
<tr>
<th>Dose</th>
<th>Description</th>
</tr>
</thead>
<tbody>
<tr>
<td>98 nSv</td>
<td>banana equivalent dose, a whimsical unit of radiation</td>
</tr>
<tr>
<td>0.25 μSv</td>
<td>U.S. limit on effective dose from a single airport security screening</td>
</tr>
<tr>
<td>5 to 10 μSv</td>
<td>one set of dental X-rays</td>
</tr>
<tr>
<td>80 μSv</td>
<td>average dose to people living within 16 km of Three Mile Island accident</td>
</tr>
<tr>
<td>0.4 to 0.6 mSv</td>
<td>two-view mammogram, using weighting factors updated in 2007</td>
</tr>
<tr>
<td>2 to 7 mSv</td>
<td>barium fluoroscopy</td>
</tr>
<tr>
<td>10 to 30 mSv</td>
<td>single full-body CT scan</td>
</tr>
<tr>
<td>68 mSv</td>
<td>estimated maximum dose to evacuees who lived closest to the Fukushima I nuclear accidents</td>
</tr>
<tr>
<td>0.67 Sv</td>
<td>highest dose received by a worker responding to the Fukushima emergency</td>
</tr>
</tbody>
</table>

### Chronic Dose Examples

<table>
<thead>
<tr>
<th>Dose</th>
<th>Description</th>
</tr>
</thead>
<tbody>
<tr>
<td>1 mSv/yr</td>
<td>ICRP recommended maximum for artificial irradiation of the public, excluding medical and occupational exposures.</td>
</tr>
<tr>
<td>2.4 mSv/yr</td>
<td>Natural background radiation, global average</td>
</tr>
<tr>
<td>24 mSv/yr</td>
<td>Natural background radiation at airline cruise altitude</td>
</tr>
<tr>
<td>9 Sv/yr</td>
<td>NRC definition of a high radiation area in a nuclear power plant, warranting a chain-link fence</td>
</tr>
<tr>
<td>&gt;90 kSv/yr</td>
<td>most radioactive hotspot found in Fukushima I in areas normally accessible to workers</td>
</tr>
<tr>
<td>2.3 MSv/yr</td>
<td>typical PWR spent fuel bundle, after 10 year cool down, no shielding</td>
</tr>
</tbody>
</table>

Nuclear Stability and Binding Energy

- **Binding energy** = the amount of energy released when a nucleus is formed.
- When two protons and two neutrons combine to form a helium nucleus, energy is released. This is the total binding energy for the helium nucleus.
The **binding energy per nucleon**, which is the total binding energy divided by the number of nucleons (protons and neutrons), is a good indication of nuclear stability.

For example, because a uranium-235 atom has many more nucleons than an iron-56 atom, it has a much larger total binding energy, but an iron-56 atom is significantly more stable than a uranium-235 atom. This is reflected in the higher binding energy per nucleon for iron-56.

Binding energy per nucleon generally increases from small atoms to atoms with a mass number around 56.

Binding energy per nucleon generally decreases from atoms with a mass number around 56 to larger atoms.
Because binding energy per nucleon generally increases from small atoms to atoms with a mass number around 56, fusing small atoms to form larger atoms (nuclear fusion) releases energy.

Because binding energy per nucleon generally decreases from atoms with a mass number around 56 to larger atoms, splitting large atoms to form medium-sized atoms (nuclear fission) also releases energy.
Nuclear Fusion

\[
\frac{2}{1}H + \frac{3}{1}H \rightarrow \frac{4}{2}He + \frac{1}{0}n
\]

Deuterium + Tritium \rightarrow Helium + Neutron

- Products are much more stable than reactants, so products have much lower PE, and a lot of energy is released.
Nuclear Fusion

\[ \frac{2}{1}H + \frac{3}{1}H \rightarrow \frac{4}{2}He + \frac{1}{0}n \]

Deuterium + Tritium \( \rightarrow \) Helium + Neutron

- Requires a very high temperature (about \(10^6 \, ^\circ C\)) to initiate the fusion.
  - The electromagnetic repulsion between the positive nuclei is felt at a relatively long range.
  - The strong force attraction is only significant when the nuclei are very close.
  - Therefore, unless the nuclei are rushing together at a very high velocity (very high temperature), the +/+ repulsion slows the nuclei down, stops them, and accelerates them away from each other before they are close enough for the strong force to play a role.
Nuclear Fusion
Powers the Sun

\[
\begin{align*}
\frac{1}{1}H + \frac{1}{1}H & \rightarrow \frac{2}{1}H + \frac{0}{+1}e \\
\frac{2}{1}H + \frac{1}{1}H & \rightarrow \frac{3}{2}He \\
\frac{3}{2}He + \frac{3}{2}He & \rightarrow \frac{4}{2}He + \frac{1}{1}H + \frac{1}{1}H
\end{align*}
\]
Nuclear Fission

\[ ^{235}_{92}U + ^{1}_{0}n \rightarrow ^{236}_{92}U + ^{138}_{56}Ba + 3 ^{1}_{0}n \rightarrow \text{Energy} \]
Chain Reaction

$^{235}_{92}U$ → $^{235}_{92}U$ → fission products

$^{235}_{92}U$ → $^{235}_{92}U$ → fission products

$^{235}_{92}U$ → $^{235}_{92}U$ → fission products

$^{235}_{92}U$ → $^{235}_{92}U$ → fission products
Nuclear Power Plant

Pressurized Water Reactor (PWR)
Nuclear Power Plant

Boiling Water Reactor (BWR)
To get a sustained chain reaction, the percentage of $^{235}\text{U}$ must be increased to about 3%, in part because the unfissionable $^{238}\text{U}$ absorbs too many neutrons.

\[
\begin{align*}
^{238}\text{U} & + \frac{1}{0}\text{n} \rightarrow ^{239}\text{U} \\
^{239}\text{U} & \rightarrow ^{239}\text{Np} + ^{0}\text{e} \\
^{239}\text{Np} & \rightarrow ^{239}\text{Pu} + ^{0}\text{e}
\end{align*}
\]
• Fuel rods
  – A typical 1000-megawatt power plant will have from 90,000 to 100,000 kg of enriched fuel packed in 100 to 200 zirconium rods about 4 meters long.

• Moderator slows neutrons
  – $^{235}$U atoms are more likely to absorb slow neutrons.
  – Can be water
Nuclear Power Plant

• Control Rods
  – Substances, such as cadmium or boron, that absorb neutrons.
  – Control rate of chain reaction
  – Dropped at first sign of trouble to stop fission reaction

• Facility includes instrumentation to monitor and control the reactor, radiation shielding, and a containment building.
Nuclear Power Plant

- Fission reactions provide heat, which is used to boil water to create steam, which turns a steam turbine to generate electricity.

- Get heat from
  - Fission reaction
  - Radioactive decay of fission products
  - Gamma rays released converted into heat
Thermal Reactors

• Thermal reactors use slowed or thermal neutrons.
• Almost all current reactors are of this type.
• A moderator slows neutrons until their kinetic energy approaches the average kinetic energy of the surrounding particles.
  • Moderator can be regular (light) water (74.8% of the world's reactors), solid graphite (20% of reactors) and heavy water (dueterium oxide - 5% of reactors).
• Thermal neutrons have a far higher probability of fissioning the fissile nuclei (\(^{235}\text{U}, \, ^{239}\text{Pu}, \text{ and } ^{241}\text{Pu}\)), and a relatively lower probability of neutron capture by \(^{238}\text{U}\) compared to the faster neutrons that originally result from fission, allowing use of low-enriched uranium fuel. The moderator is often also the coolant, usually water under high pressure to increase the boiling point.
Fast Neutron Reactors

- Use fast neutrons to cause fission in their fuel.
- Do not have a neutron moderator, and use less-moderating coolants.
- Requires the fuel to be more highly enriched in fissile material (about 20% or more) due to the relatively lower probability of fission versus capture by U-238.
- More difficult to build and more expensive to operate.
- Less common than thermal reactors in most applications.
Diablo Canyon Nuclear Plant

• At Avila Beach in San Luis Obispo County, California.
• Two 1,100 MWe (MW of electrical energy produced) pressurized-water nuclear reactors operated by Pacific Gas & Electric produce about 18,000 GW·h of electricity annually, supplying the electrical needs of more than 2.2 million people.
**Uranium Fuel Cycle**

**Uranium** is mined, enriched and manufactured to make nuclear fuel (1), which is delivered to a nuclear power plant. After usage in the power plant the spent fuel might be delivered to a reprocessing plant (if fuel is recycled) (2) or to a final repository (if no recycling is done) (3) for geological disposition. In reprocessing, 95% of spent fuel can be recycled to be returned to usage in a nuclear power plant (4).

http://en.wikipedia.org/wiki/Nuclear_fuel_cycle
Uranium Ore

• Uranium is one of the more common elements in the Earth's crust, some 40 times more common than silver and 500 times more common than gold.

• The primary uranium ore is uraninite, which is largely $\text{UO}_2$, but also contains $\text{UO}_3$ and oxides of other metals. It is commonly known as pitchblende.
Uranium Enrichment

- Uranium in uranium ore is about 99.3% $^{238}\text{U}$, which is not fissionable, and 0.7% $^{235}\text{U}$, which is fissionable.
- Needs to be enriched in $^{235}\text{U}$ to varying degrees (depending on the reactor design) to be useful in a nuclear power plant.

Highly enriched uranium metal
Uranium Enrichment

- In a series of steps, the uranium in uranium ore is converted into uranium hexafluoride, UF$_6$.
- It sublimes (goes directly from solid to gas) at 56.5 °C.
- The UF$_6$ can be enriched in $^{235}$UF$_6$ by either gas diffusion (first generation) or using gas centrifuges (second generation), which requires less energy.
Gas Centrifuge

- Creates a strong centrifugal force so that the heavier gas molecules containing $^{238}\text{U}$ move toward the outside of the cylinder and the lighter gas molecules rich in $^{235}\text{U}$ collect closer to the center.
- A large number of rotating cylinders connected in series and parallel formations.

The bottom of the rotating cylinder can be heated, producing convection currents that move the $^{235}\text{U}$ up the cylinder, where it can be collected.
Depleted Uranium

- Most of the depleted uranium produced is stored as uranium hexafluoride, DUF$_6$, in steel cylinders in open air yards close to enrichment plants.
Spent Fuel Rods

- Fission products that emit beta and gamma radiation
- Some fissionable U-235 and Pu-239
- Alpha emitters, such as uranium-234, neptunium-237, plutonium-238 and americium-241
- Sometimes some neutron emitters such as californium (Cf).
Nuclear Reprocessing

• Process to chemically separate and recover fissionable plutonium and uranium from irradiated nuclear fuel.

• Purposes
  – Originally reprocessing was used solely to extract plutonium for producing nuclear weapons.
  – The reprocessed plutonium can be recycled back into fuel for nuclear reactors.
  – The reprocessed uranium, which constitutes the bulk of the spent fuel material, can in principle also be re-used as fuel, but that is only economic when uranium prices are high.

http://en.wikipedia.org/wiki/Nuclear_reprocessing
Nuclear Reprocessing

- Reprocessing of civilian fuel has long been employed in France, the United Kingdom, Russia, Japan, and India.
- Briefly done at the West Valley Reprocessing Plant in the United States.
- In October 1976, concerned about nuclear weapons proliferation, President Gerald Ford indefinitely suspended the commercial reprocessing and recycling of plutonium in the U.S.
- In March 1999, the U.S. Department of Energy (DOE) reversed its policy and signed a contract with a consortium to design and operate a mixed oxide (MOX) fuel fabrication facility. There are no customers yet.

http://en.wikipedia.org/wiki/Nuclear_reprocessing
Liquid-Liquid Extraction

Polar compounds will congregate in “aqueous” layer

Non-polar compounds will congregate in “organic” layer

Typically performed in a separatory funnel:

Aqueous layer (polar things)

Organic layer (non-polar things)
PUREX Process

- Dissolve in 7 M HNO₃.
- Filter out solids
- Combine with 30% tributyl phosphate (TBP) to form \( \text{UO}_2(\text{NO}_3)_2 \cdot 2\text{TBP} \) and \( \text{PuO}_2(\text{NO}_3)_2 \cdot 2\text{TBP} \) complexes.
- Extract with an organic solvent, such as kerosene.
  - \( \text{UO}_2(\text{NO}_3)_2 \cdot 2\text{TBP} \) and \( \text{PuO}_2(\text{NO}_3)_2 \cdot 2\text{TBP} \) complexes in nonpolar organic solvent
  - Fission products, and transuranium elements americium and curium remain in the aqueous phase.

http://en.wikipedia.org/wiki/PUREX
Separation of U, Pu, and Fission Products

- **Nonpolar solvent Layer**
- **Aqueous Layer**

**On to Uraniun purification**

**On to Plutonium purification**

**Fission products, minor actinides:** discarded as "high level waste"
• Plutonium is separated from uranium in a separate extraction by treating the kerosene solution with aqueous iron(II) sulfamate, Fe(SO$_3$NH$_2$)$_2$, which reduces the plutonium to the +3 oxidation state. The plutonium passes into the aqueous phase.

• Variations on the PUREX process have been developed.

http://en.wikipedia.org/wiki/PUREX
• 2002 – China shipped about 20 tons of tributyl phosphate (TBP) to North Korea.
• Considered to be sufficient to extract enough material for three to five nuclear weapons
Nuclear Waste Storage and Disposal

- Nuclides of special concern
  - Tc-99 (half-life 220,000 years) and I-129 (half-life 17 million years), which dominate spent fuel radioactivity after a few thousand years.
  - Np-237 (half-life two million years) and Pu-239 (half-life 24,000 years).
- Needs treatment, followed by a long-term management strategy involving storage, disposal, or transformation of the waste into a non-toxic form.
- Governments around the world are considering a range of waste management and disposal options, though there has been limited progress toward long-term waste management solutions.

Nuclear Waste Storage/Disposal Possibilities

- "Long term above ground storage", not implemented.
- "Disposal in outer space", not implemented.
- "Deep borehole disposal", not implemented.
- "Rock-melting", not implemented.
- "Disposal at subduction zones", not implemented.
- "Ocean disposal", done by USSR, UK, Switzerland, USA, Belgium, France, Netherland, Japan, Sweden, Russia, Germany, Italy and South Korea. (1954–93) It's not permitted by international agreements.
- "Sub seabed disposal", not implemented, not permitted by international agreements.
- "Disposal in ice sheets", rejected in Antarctic Treaty
- "Direct injection" of liquid waste, done by USSR and USA.

Yucca Mountain Nuclear Waste Repository

- Deep geological repository storage facility for spent nuclear reactor fuel and other high level radioactive waste,
- Approved in 2002 by the United States Congress.
- Project was defunded in 2010.
- The US Government Accountability Office stated that the closure was for political, not technical or safety reasons.
Yucca Mountain Nuclear Waste Repository

- United States without any long term storage site for high level radioactive waste, currently stored on-site at various nuclear facilities around the country

http://en.wikipedia.org/wiki/Yucca_Mountain_nuclear_waste_repository
The Blue Ribbon Commission established by the Secretary of Energy released its final report on January 26, 2012. It expressed urgency to find a consolidated, geological repository, but also that any future facility should have input from the citizens around it.

http://en.wikipedia.org/wiki/Yucca_Mountain_nuclear_waste_repository
Status of Nuclear Power

- 437 nuclear power plants in operation
- 371.762 GW_e total installed capacity
- 64 Nuclear power plants under construction.

http://www.iaea.org/pris/
Total Number of Reactors: 437

United States of America: 60
France: 58
Japan: 58
Russia: 49
Korea, Republic of: 47
Canada: 27
India: 25
China: 23
United Kingdom: 18
Ukraine: 16
Sweden: 16
Germany: 15
Spain: 15
Belgium: 14
Czech Republic: 13
Switzerland: 13
Finland: 13
Hungary: 12
Slovakia: 12
Pakistan: 11
Argentina: 11
Brazil: 11
Bulgaria: 11
Mexico: 10
Romania: 10
South Africa: 9
Armenia: 9
Iran, Islamic Republic of: 9
Netherlands: 8
Slovenia: 8

Worldwide Nuclear Reactors

http://www.iaea.org/pris/
United States

- 104 nuclear power plants
- 1 under construction
- $7.9 \times 10^5$ GWh in 2011

http://www.iaea.org/pris/
Carbon emissions from nuclear power
Sovacool life cycle study survey, 2008

- Front end, 25.09 g/kWh
- Construction, 8.20 g/kWh
- Operation, 11.58 g/kWh
- Back end, 9.20 g/kWh
- Decommissioning, 12.01 g/kWh

Total, 66.08 g/kWh

Mean value of carbon dioxide emissions from qualified life cycle studies among 103 surveyed. Includes results of 1997 Vattenfall study.
# Lifecycle Greenhouse Gas Emission Estimates for Electricity Generators

<table>
<thead>
<tr>
<th>Technology</th>
<th>Description</th>
<th>Estimate (g CO2/kWh\textsubscript{e})</th>
</tr>
</thead>
<tbody>
<tr>
<td>Wind</td>
<td>2.5 MW offshore</td>
<td>9</td>
</tr>
<tr>
<td>Hydroelectric</td>
<td>3.1 MW reservoir</td>
<td>10</td>
</tr>
<tr>
<td>Wind</td>
<td>1.5 MW onshore</td>
<td>10</td>
</tr>
<tr>
<td>Biogas</td>
<td>Anaerobic digestion</td>
<td>11</td>
</tr>
<tr>
<td>Hydroelectric</td>
<td>300 kW run-of-river</td>
<td>13</td>
</tr>
<tr>
<td>Solar thermal</td>
<td>80 MW parabolic trough</td>
<td>13</td>
</tr>
<tr>
<td>Biomass</td>
<td>various</td>
<td>14-35</td>
</tr>
<tr>
<td>Solar PV</td>
<td>Polycrystaline silicon</td>
<td>32</td>
</tr>
<tr>
<td>Geothermal</td>
<td>80 MW hot dry rock</td>
<td>38</td>
</tr>
<tr>
<td>Nuclear</td>
<td>various reactor types</td>
<td>66</td>
</tr>
<tr>
<td>Natural gas</td>
<td>various combined cycle turbines</td>
<td>443</td>
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<tr>
<td>Fuel Cell</td>
<td>hydrogen from gas reforming</td>
<td>664</td>
</tr>
<tr>
<td>Diesel</td>
<td>various generator and turbine types</td>
<td>778</td>
</tr>
<tr>
<td>Heavy oil</td>
<td>various generator and turbine types</td>
<td>778</td>
</tr>
<tr>
<td>Coal</td>
<td>various generator types with scrubbing</td>
<td>960</td>
</tr>
<tr>
<td>Coal</td>
<td>various generator types without scrubbing</td>
<td>1050</td>
</tr>
</tbody>
</table>

Retiring Nuclear Power Plants

Impact of NPP Retirement on Carbon Emissions

*Data does not represent future growth in demand but only represents the increase in CO2 emissions to replace what nuclear produces today.*